

2,11-DISELENA[3,3]METACYCLOPHANE - SYNTHESIS
AND VARIABLE TEMPERATURE PMR SPECTRUM

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Despite the considerable recent interest¹⁻³ in dithia[3,3]cyclophanes no diselena[3,3]cyclophane has thus far been reported. I now report the synthesis and variable temperature pmr spectrum of 2,11-diselena[3,3]metacyclophane **1**. Reaction of *m*-xylylene bromide with anhydrous sodium selenide (Alfa) in dry DMSO or hexamethylphosphoramide gave 5-10% yields of **1**, mp 121-122°. Use of aqueous or ethanolic solvents or attempted isolation of selenols gave heavy contamination with diselenides.^{4,5} The structure of **1** was confirmed by analysis, mass (M^+ at *m/e* 368) and pmr spectroscopy (figure 1). At 20°C the methylene bridge pmr absorptions (3.78 δ) are deshielded by 0.24 ppm, whereas the internal aromatic hydrogens (6.46 δ) are shielded by 0.58 ppm from those of the cyclic trimer (mp 130-132°). This effect is more marked than in the case of 2,11-dithia[3,3]metacyclophane.³ The room temperature spectrum is consistent with the molecule conformationally flipping between various syn and anti conformers.¹ Cooling to -110°C freezes out one particular anti conformer (such as that shown in the photograph) in which one of the internal protons is pushed directly into the opposite π cloud and is highly shielded to 4.50 δ , the other internal proton being in a normal aromatic environment. The average position of these two protons would be ca. 5.7 δ if a simple exchange process were occurring at room temperature. The observed average position indicates that other conformers have become more dominant. From the coalescence temperature of -95°C and $\Delta\nu = 226\text{Hz}$ an estimate⁷ of the energy barrier to flipping, ΔG^\ddagger is 8.0 kcal/mole. This presumably reflects the slightly greater size of selenium compared with sulfur, since the corresponding thiacyclophane is still mobile at -110°C. We are currently investigating the chemistry of this compound and thank the University of Victoria and National Research Council for support of this work.

REFERENCES

1. T. Sato, M. Wakabayashi, K Hata and M. Kainosho, *Tetrahedron*, **27**, 2737 (1971); F. Vögtle and L. Schunder, *Chem. Ber.* **102**, 2677 (1969).
2. V. Boekelheide, I.D. Reingold, and M. Tuttle, *J.C.S. Chem. Comm.*, 406 (1973); N. Kannen, T. Umemoto, T. Otsubo and S. Misumi, *Tetrahedron Letters*, 4537 (1973).
3. R.H. Mitchell and V. Boekelheide, *J. Am. Chem. Soc.*, **96**, 1547 (1974) and accompanying papers.
4. W.H.H. Günther and M.N. Salzman, *Ann. N.Y. Acad. Sci.*, **192**, 25 (1972).
5. 2,3,12,13-tetraselena[4,4]metacyclophane previously claimed by Günther⁴ has been prepared⁶ by alkaline hydrolysis of m-xylylene selenocyanate and converted into **1** by action of tris-diethylaminophosphine in benzene.
6. R. H. Mitchell, submitted to *Can. J. Chem.*
7. I.C. Calder and P.J. Garratt, *J. Chem. Soc. (B)*, 660 (1967).

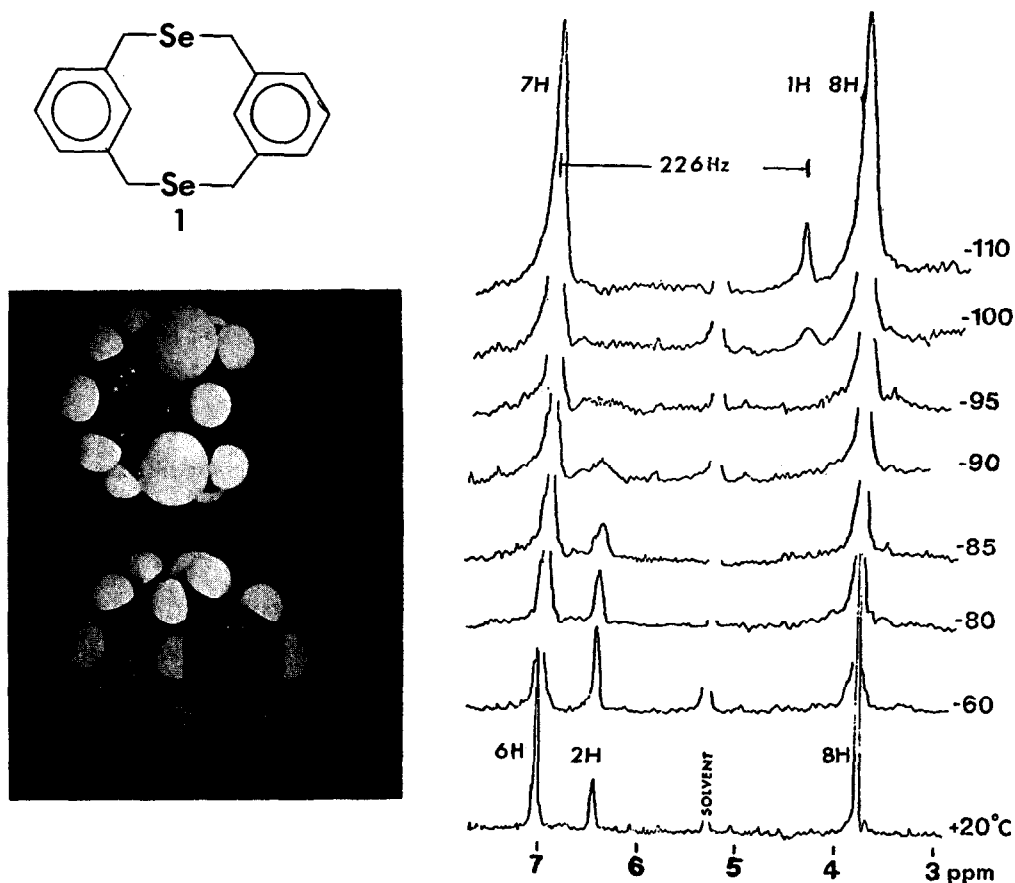


Figure 1 - Pmr spectrum of **1** in $\text{CD}_2\text{Cl}_2\text{-CDCl}_3$ (3:1) recorded on a Perkin-Elmer R32-90MHz spectrometer.